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Article

# Theoretical Study on Non-Linear Optics Properties of Polycyclic Aromatic Hydrocarbons and the Effect of Their Intercalation with Carbon Nanotubes

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**Abstract:** Results of a theoretical study devoted to comparing NLO (non-linear optics) responses of derivatives of tetracene, isochrysene, and pyrene are reported. The static hyperpolarizability  $\beta$ , the dipole moment  $\mu$ , the HOMO and LUMO orbitals, and their energy gap were calculated using the CAM-B3LYP density functional combined with the cc-pVDZ basis set. The para-disubstituted NO<sub>2</sub>-tetracene-N(CH<sub>3</sub>)<sub>2</sub> has the highest NLO response, which is related to a large intramolecular charge transfer. Adding vinyl groups to the para-disubstituted NO<sub>2</sub>-tetracene-N(CH<sub>3</sub>)<sub>2</sub> results in an increase in the NLO responses. We further investigated the effect of the intercalation of various push–pull molecules inside an armchair single-walled carbon nanotube. The intercalation leads to increased NLO responses, something that depends critically on the position of the guest molecule and/or on functionalization of the nanotube by donor and attractor groups.

**Keywords:** conjugated molecules; non-linear optical responses; functionalization; embedding



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## 1. Introduction

The backbone of polycyclic aromatic hydrocarbons (PAHs) contains a sequence of at least two fused benzene rings whereby the way they are linked distinguishes different PAHs [1,2]. PAHs can have an unlimited number of contiguous rings [3–5]. This gives rise to a large number of isomers and enriches this family of aromatic hydrocarbons. The main approach for producing PAHs is through an incomplete combustion of organic materials (for instance, fuels and coal) [6–8]. PAHs are divided into two classes (light and heavy) according to the number of rings involved in their structures. Each class has its own physicochemical properties [9], which allows for a large variety of different applications including organic field effect transistors [10–12], organic light-emitting diodes [13], reinforcing agents in pigment lasers [14], and batteries [15].

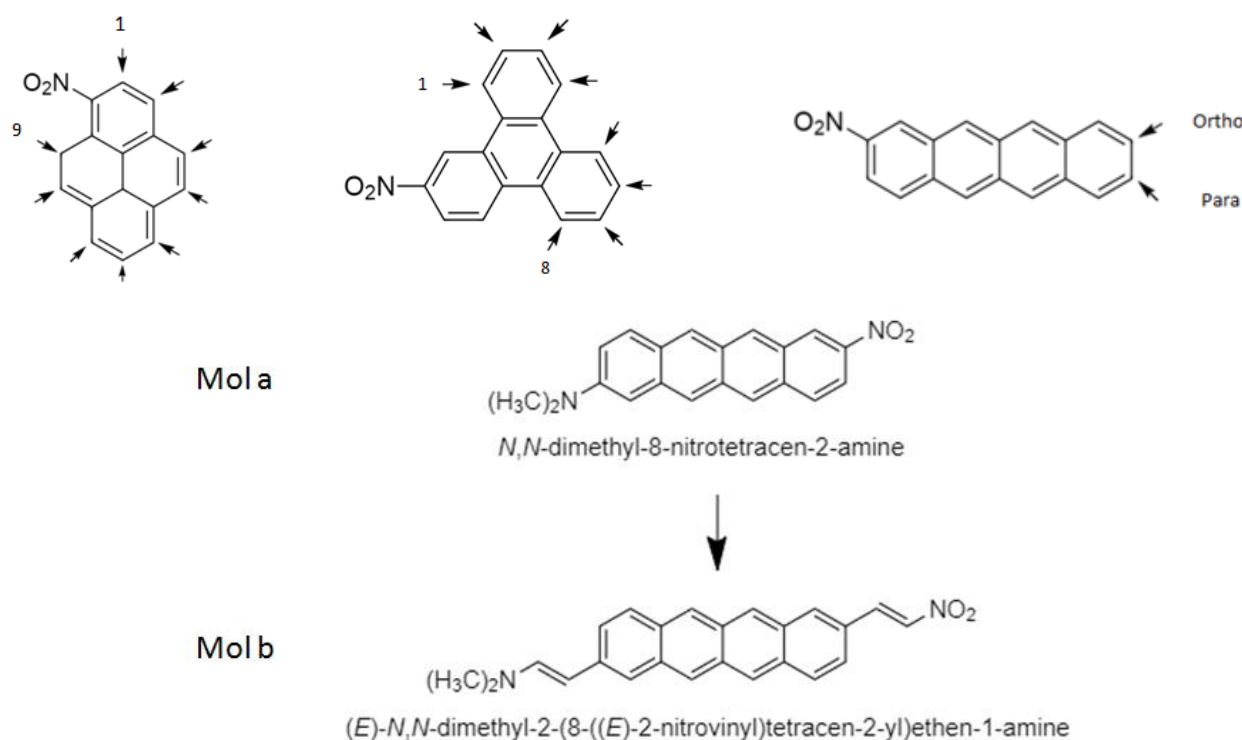
The aim of the present study is to use theoretical methods to study the performance of three smaller PAHs, i.e., tetracene, isochrysene (or triphenylene), and pyrene, with special emphasis on their non-linear optics (NLO) responses. The  $\pi$  electrons of these conjugated molecules [16–18] facilitate an intramolecular charge transfer (ICT) between electron donor (D) and electron acceptor (A) groups when such groups are attached [19,20]. To study how the NLO responses can be influenced upon functionalization of the system is one purpose of the present work.

A number of recent papers have focused on intramolecular charge transfer in PAHs, including studies on tetracyclic molecules and their derivatives [21–25]. Moreover, it has been shown that purely organic rings can be considered as being more aromatic than BN-containing systems [26] and, accordingly, to have more delocalized  $\pi$  electrons.

Even carbon nanotubes (CNTs) can be considered as being a special case of extended PAHs, independent of whether they are single-walled carbon nanotubes, SWCNTs; or multi-walled carbon nanotubes, MWCNTs. Since their discovery in 1991 [27], a vast number of studies of their properties have appeared, including studies of their practical applications in, e.g., pharmacy, mechanics, and optoelectronics. They possess a high mechanical resistance, a high electrical and thermal conductivity, and chemical inertness [28–30]. Because of their optoelectronic properties, they have been used for light-emitting diodes [31]. In addition, functionalization of SWCNTs has been used as a way of improving their properties as shown, e.g., by Khazaei et al. [32]. The hollow structure of the carbon nanotubes, shared by the fullerenes, allows for intercalation, a prospect that has been studied by, e.g., Hirscher et al. [33] and by Chaban et al. [34,35].

In addition, NLO properties of such systems have been at the center of earlier studies [36,37]. However, a more systematic study of the dependence of the NLO properties on the size of the system, on functionalization, and on intercalation is lacking, although this could provide very useful information for experimentalists who aim at designing optimal systems. It is the purpose of the present work to provide results of such a study.

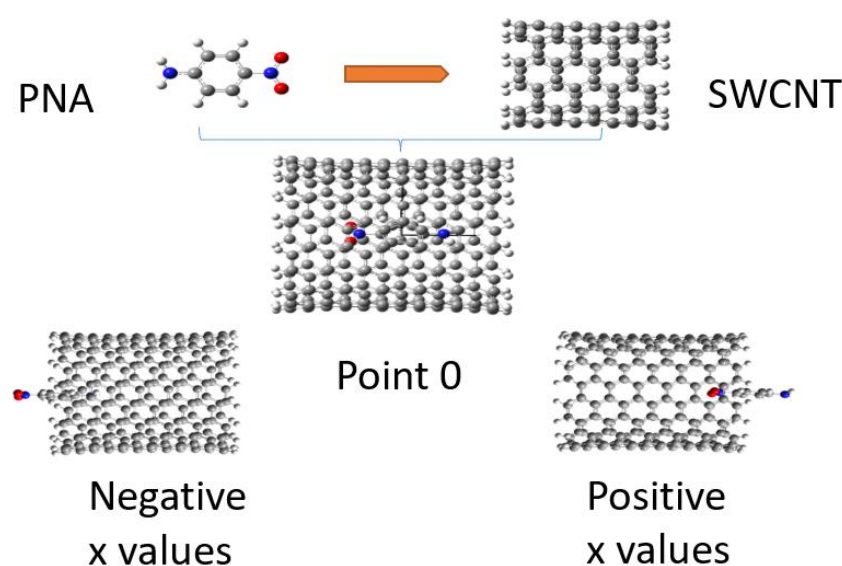
We also study some push–pull molecules when interacting with SWCNTs. The push–pull molecules considered in this work are shown on Figure 1. They consist of the pure polycyclic aromatic hydrocarbons tetracene, isochrysene, and pyrene and they all contain a conjugated bridge with delocalized  $\pi$  electrons [38–40]. We study the effects of substituting the PAHs at different positions through donors [41]:  $\text{NH}_2$ ,  $\text{N}_2\text{H}_3$ ,  $\text{N}(\text{CH}_3)_2$ ,  $\text{OH}$ ,  $\text{OCH}_3$ , and the acceptor group  $\text{NO}_2$  [42].



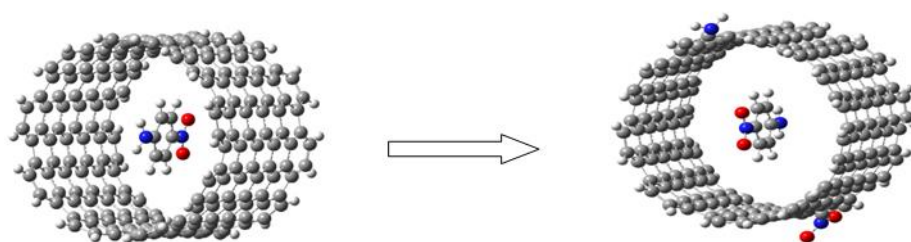
**Figure 1.** Schematic representation of the molecules of the present study. The arrows show positions at which functional groups are attached.

For the molecule giving the largest NLO response, i.e., tetracene, we subsequently studied modified versions of this containing a larger conjugated part. This was achieved by the addition of vinyl groups at the terminations of tetracene. Then, we compared the basic molecule (Mol a) and the derivative obtained after the modification (Mol b) in terms of intramolecular charge transfer, the first hyperpolarizability, and the dipole moment.

Subsequently, we considered the effects of intercalation of derivatives of a single PAH molecule inside carbon nanotubes. Initially, we constructed a (9,9) armchair nanotube with a diameter of 12.21 nm and a length of 19.69 nm as shown, e.g., in Figures 2 and 3. Dangling bonds at the ends were saturated with hydrogen atoms. The initial structure of this chair-type SWCNT was optimized using the B3LYP density functional [43] combined with the 6-31g(d,p) basis set [44]. We investigated the effect of the position of the guest molecule, paranitroaniline (PNA), inside the nanotube by performing single-point calculations using the CAM-B3LYP functional together with the GD3 dispersion correction [45,46] and using the 6-31g(d,p) basis set. Various NLO parameters, including the static first hyperpolarizability, the dipole moment, and the HOMO–LUMO energy gap, were calculated for different positions of the guest molecule inside the nanotube by translating the former along the x-axis (parallel to the tube) with a step length of 2 Å relative to the initial position (denoted position 0, cf., Figure 2).

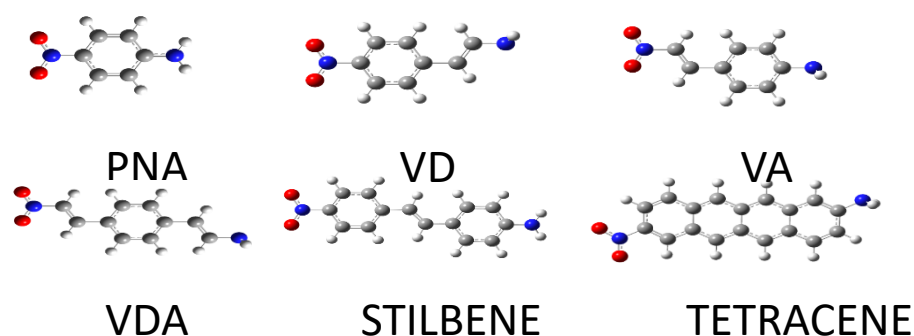


**Figure 2.** Structures showing the positions of the guest molecule (PNA) inside the host molecule (SWCNT).



**Figure 3.** Armchair-type nanotube structure to which a donor and an electron acceptor group have been attached.

After that, we examined the effect of the size of the guest molecule on the intramolecular charge transfer of the system. For that purpose, we considered different push–pull molecules inserted inside the chair-like nanotube. As guest molecules, we considered PNA, VD, VA, VDA, stilbene, and tetracene, all shown in Figure 4. For each system, the static hyperpolarizabilities and the dipole moment were calculated



**Figure 4.** Structures of the push–pull molecules that were inserted inside the armchair-type carbon nanotube.

Finally, we modified the host system, i.e., to the armchair-type nanotube, we attached an  $\text{NH}_3$  donor on one side and an  $\text{NO}_2$  acceptor group on the other side, cf., Figure 3. At first, the structure of the isolated host was optimized using B3LYP/6-31g(d,p), after which the push–pull molecule PNA was inserted in the center of the tube and calculations were performed to check the effect of these substitutions on the hyperpolarizabilities and on the total dipole moment.

## 2. Computational Details

At first, we emphasize that our study involves several approximations. The size and number of the systems of our interest make it prohibitive to apply the most accurate computational methods for each of those. Instead, our focus is on studying the changes when modifying the systems in one way or another, so that our results should be able to describe those changes, although the absolute numbers will be less accurate. The approximations we employ include a basis set of finite size, the finite lengths of the carbon nanotubes, and the density functional itself.

All structures were visualized using the Gaussview 5.0 software [47] and all calculations were performed using Gaussian 09 [48]. Geometry optimizations of all the molecules and their derivatives were carried out using the density functional B3LYP [49–51] combined with the cc-pVDZ basis set [52–54].

Initially, we performed a benchmark study comparing static hyperpolarizabilities  $\beta_{tot}$  obtained using a number of hybrid functionals such as PBE0, BMK, BHHLYP, M062X, and CAM-B3LYP in which the Hartree–Fock exchange is partly incorporated, i.e., to 25%, 42%, 50%, 54%, and 65%, respectively [55–59]. The results were compared to those obtained using the MP2 method (the second-order Møller–Plesset perturbation method). The latter was taken as a reference due to the absence of experimental values as MP2 results often are considered accurate [60–63]. This comparison was performed only on tetracene derivatives in order to determine the appropriate functional for this kind of system and for the first static hyperpolarizability  $\beta$  (cf., Equation (1)) [64–66].

In the general case, the first hyperpolarizability is a  $3 \times 3 \times 3$  tensor that, however, can be reduced to 10 numbers with the help of the Kleinman notation,  $\beta_{xxx}$ ,  $\beta_{xxy}$ ,  $\beta_{xyy}$ ,  $\beta_{yyy}$ ,  $\beta_{xxz}$ ,  $\beta_{xyz}$ ,  $\beta_{yyz}$ ,  $\beta_{xzz}$ ,  $\beta_{yzz}$ , and  $\beta_{zzz}$  [67].

We focused on the total hyperpolarizability:

$$\beta_{tot} = \left( \beta_x^2 + \beta_y^2 + \beta_z^2 \right)^{\frac{1}{2}} \quad (1)$$

with

$$\beta_x = (\beta_{xxx} + \beta_{xyy} + \beta_{xzz}) \quad (2)$$

$$\beta_y = (\beta_{yyy} + \beta_{yzz} + \beta_{yxx}) \quad (3)$$

$$\beta_z = (\beta_{zzz} + \beta_{zyy} + \beta_{zxx}) \quad (4)$$

According to our benchmark study, the CAM-B3LYP functional provides the best agreement with the MP2 reference results. Therefore, this functional was used in the subsequent calculations. This finding agrees with that of Rabah et al. [52].

Subsequently, we performed single-point (SP) calculations using the CAM-B3LYP functional combined with the cc-pVDZ basis set on each molecule. This functional includes a description of long-range corrections [68,69] and, accordingly, it provides a better description of properties related to an intramolecular charge transfer [70].

The dipole moment was calculated according to [71,72]

$$\mu = \sqrt{(\mu_x^2 + \mu_y^2 + \mu_z^2)} \quad (5)$$

We also used the energy gap between the HOMO and the LUMO frontier orbitals:

$$\Delta E_{H-L} = \varepsilon_{LUMO} - \varepsilon_{HOMO} \quad (6)$$

as parameters quantifying the NLO properties of our systems.

For the geometric structure, we focused on the BLA (Bond Length Alternation) parameter, i.e., the difference between the average lengths of single and double bonds in a conjugated system [73]. A smaller value of the BLA facilitates an intramolecular charge transfer.

### 3. Results and Discussion

#### 3.1. I-NLO Responses of PAHs

##### 3.1.1. Ia. Selection of the Functional

In this part, we identify the density functional that gives results closest to those obtained with the MP2 method. The latter is considered as reliable for NLO properties. DFT (functionals BMK, BHHLYP, CAM-B3LYP, M062X, and PBE0) as well as MP2 calculations were performed in combination with the cc-pVDZ basis set to calculate the first hyperpolarizabilities of ten tetracene derivatives.

The results (see Table 1) show that the PBE0 functional overestimates the hyperpolarizabilities. The values related to the functionals BMK, BHHLYP, and M062X give a less pronounced difference, whereas the best agreement is obtained for the functional CAM-B3LYP. This is explained by the fact that this functional includes long-range Hartree–Fock exchange interactions. Consequently, the subsequent calculations for the pyrene and isochrysene derivatives were carried out using this functional in combination with the cc-pVDZ basis set. That this combination yields accurate results, particularly concerning trends, is in agreement with our earlier findings [52].

**Table 1.** First hyperpolarizability of tetracene derivatives (in  $10^{-30}$  esu) using different functionals with the cc-pVDZ basis set (O: Ortho; P: Para positions of substitutions).

	NH2 -O	NH2 -P	N(CH3)2 -O	N(CH3)2 -P	OH -O	OH -P	OCH3 -O	OCH3 -P	N2H3 -O	N2H3 -P
BMK	46.71	34.64	65.77	11.01	35.68	49.82	40.66	58.78	54.94	84.75
BHHLYP	41.89	32.22	56.49	92.62	32.58	44.90	36.40	51.62	48.09	73.02
CAM-B3LYP	39.95	28.55	54.45	85.07	30.62	41.38	34.52	47.83	46.31	67.71
PBE0	84.29	67.14	120.27	235.47	67.49	104.58	78.81	129.36	100.44	179.01
M06-2X	40.68	28.25	57.43	25.78	30.28	41.22	34.49	48.14	47.84	70.89
MP2	45.65	29.78	63.76	91.29	34.06	42.51	38.57	49.16	52.83	71.22

##### 3.1.2. Ib. Study of Intramolecular Charge Transfer

The calculated static first-order hyperpolarizabilities reported in Table 2 show that among the tetracene derivatives, the para-disubstituted NO<sub>2</sub>-tetracene-N(CH<sub>3</sub>)<sub>2</sub> gives the largest value of  $\beta$  as well as the largest dipole moment  $\mu$ , and also the lowest-energy gap, which is roughly inversely proportional to an intramolecular charge transfer. It is

added that a comparison of the dipole moment or the hyperpolarizability between different molecules is hampered by the fact that these properties are extensive properties, so, in general, larger molecules have larger values for these properties. However, the differences we discuss here are larger than what can be explained through this simple fact.

**Table 2.** Calculated hyperpolarizability ( $10^{-30}$  esu), dipole moment (Debye), and energy gap (eV) of tetracene derivatives with substitutions at the ortho and para positions using CAM-B3LYP/cc-PVDZ. The system with the largest value for  $\beta_{tot}$  is highlighted.

	NH2-O	NH2-P	N(CH <sub>3</sub> ) <sub>2</sub> -O	N(CH <sub>3</sub> ) <sub>2</sub> -P	OH-O	OH-P	OCH <sub>3</sub> -O	OCH <sub>3</sub> -P	N2H3-O	N2H3-P
$\beta_x$	39.89	−28.52	54.42	−85.01	30.60	41.29	34.52	47.78	46.28	−67.63
$\beta_y$	2.06	−1.41	1.63	−3.02	0.93	−2.61	0.38	−2.30	1.13	−3.12
$\beta_z$	0.68	0.01	−0.62	−0.49	0.00	0.00	0.00	0.00	−1.05	−1.12
$\beta_{tot}$	39.95	28.55	54.45	85.07	30.61	41.38	34.52	47.83	46.31	67.71
$\mu_{tot}$	8.34	6.53	9.21	9.72	7.68	7.53	7.02	7.62	9.19	10.07
gap	4.61	4.76	4.50	4.50	4.66	4.68	4.60	4.62	4.58	4.61

We studied all isochrysene derivatives containing the NO<sub>2</sub> group at one side of the chromophore and an electron donor (i.e., NH<sub>2</sub>, N(CH<sub>3</sub>)<sub>2</sub>, N<sub>2</sub>H<sub>3</sub>, OH, or OCH<sub>3</sub>) at the other side (cf., Figure 1). In Table 3, we present only the results of the NLO parameters of the derivatives in which the position of the donor N(CH<sub>3</sub>)<sub>2</sub> was varied while that of the NO<sub>2</sub> group was kept fixed. This combination results in a larger ICT compared to the other combinations.

**Table 3.** Calculated hyperpolarizability ( $10^{-30}$  esu), dipole moment (Debye), and energy gap (eV) of the molecule N(CH<sub>3</sub>)<sub>2</sub>-isochrysene-NO<sub>2</sub> by varying the position of N(CH<sub>3</sub>)<sub>2</sub> from 1 to 8 (denoted mol 1 to mol 8) relative to the position of NO<sub>2</sub>, as obtained using CAM-B3LYP/cc-PVDZ. The system with the largest value for  $\beta_{tot}$  is highlighted.

	mol 1	mol 2	mol 3	mol 4	mol 5	mol 6	mol 7	mol 8
$\beta_x$	−8.53	−4.81	−9.17	10.17	−9.91	13.21	9.65	−7.32
$\beta_y$	−1.15	−7.96	−6.66	−4.72	−0.38	−3.05	−4.28	−4.05
$\beta_z$	0.00	0.00	0.00	0.00	0.00	0.00	0.00	0.00
$\beta_{tot}$	8.60	9.29	11.32	11.21	9.91	13.55	10.55	8.36
$\mu_{tot}$	5.10	4.81	8.14	7.23	7.85	5.63	7.27	7.10
gap	6.48	6.34	6.35	6.38	6.33	6.39	6.33	6.40

From Table 3, we can observe that the charge transfer occurs mainly along the x-axis (the main axis of the chromophore). Indeed, the value of  $\beta_y$  is very small compared to the value of  $\beta_x$ , and the value of  $\beta_z$  vanishes. For substitutions at positions I-6, the largest charge transfer is obtained as the donor and the acceptor groups are parallel to the dipole moment (x-axis).

Table 4 reports results obtained for pyrene derivatives substituted with N(CH<sub>3</sub>)<sub>2</sub> as a donor and NO<sub>2</sub> as an acceptor. The results are very similar to those reported in Table 3 and we, again, notice that the charge transfer occurs along the x-axis and that the substitution at positions 1-6 gives the largest charge transfer.

Table 5 summarizes the results for those derivatives of the three molecules of our interest that possess the highest values for the first hyperpolarizability. We notice that the hyperpolarizability of the tetracene derivative is markedly larger than those of the other two derivatives. The same holds for the dipole moment. The energy gap of the tetracene derivative is smaller, which correlates with the larger charge transfer between donor and acceptor.

**Table 4.** Calculated hyperpolarizability ( $10^{-30}$  esu), dipole moment (Debye), and energy gap (eV) of the molecules N(CH<sub>3</sub>)<sub>2</sub>-pyrene-NO<sub>2</sub> by varying the position from 3 to 9 of N(CH<sub>3</sub>)<sub>2</sub> relative to that of NO<sub>2</sub> (Position 1) giving mol 1-3 to mol 1-9, using CAM-B3LYP/cc-PVDZ. The system with the largest value for  $\beta_{tot}$  is highlighted.

	mol 1-3	mol 1-4	mol 1-5	mol 1-6	mol 1-7	mol 1-8	mol 1-9
$\beta_x$	3.83	−10.94	−19.48	−22.99	19.08	19.04	−10.89
$\beta_y$	9.42	−1.18	0.79	2.69	−1.44	−1.09	−1.79
$\beta_z$	0.60	0.29	0.67	0.75	0.43	0.59	0.15
$\beta_{tot}$	10.18	11.01	19.50	23.16	19.14	19.08	11.03
$\mu_{tot}$	4.92	5.91	6.65	6.89	7.62	6.20	5.02
gap	5.26	5.42	5.48	5.25	5.38	5.21	5.41

**Table 5.** Comparison between results for tetracene, isochrysene, and pyrene derivatives (first static hyperpolarizability in  $10^{-30}$  esu, dipole moment (Debye), energy gap (eV), and energies of the HOMO and LUMO orbitals in eV) as obtained using CAM-B3LYP and cc-pVDZ.

	$\beta_0$	$\mu$	gap	HOMO	LUMO
D-TETR-A	85.07	9.72	4.53	−6.13	−1.61
D-ISO-A	13.55	5.63	6.39	−7.67	−1.28
D-PYR-A	23.16	6.89	5.25	−6.75	−1.50

As the tetracene derivatives give the highest charge transfer among the three molecules, only this system is considered in the next step. In this, the  $\pi$ -conjugated system is extended by adding vinyl groups at either termination of the molecule (see Figure 1), so the effect of extending the  $\pi$ -chain length on the ICT can be analyzed [74].

This substitution leads to an increase in the first static hyperpolarizability from 85.07  $10^{-30}$  esu to 229.79  $10^{-30}$  esu. In addition, the dipole moment, which depends on the ICT, increases from 9.72 to 11.31 Debye.

For the energy gap, we notice only a smaller decrease from 4.53 eV for Mol\_a to 4.34 eV for Mol\_b. The HOMO–LUMO gap is inversely proportional to the ICT [75]. The very similar values for the gap for the two molecules can be understood from Figure 5: the frontier orbitals are largely localized to the backbone of the molecules. Equivalently, the energies of the HOMO and LUMO orbitals decrease only slightly for the substituted molecules that have a larger conjugation.

The BLA (Bond Length Alternation) parameter is useful in quantifying NLO responses for conjugated molecules. The results reported in Table 6 show an increase in BLA upon an increase in the conjugated bridge, which correlates with the previous results.

**Table 6.** Comparison between the results for tetracene and divinyl-tetracene derivatives: first static hyperpolarizability in  $10^{-30}$  esu, dipole moment (Debye), energy gap (eV), HOMO and LUMO orbital energies in eV, and BLA in Å as obtained with CAM-B3LYP and cc-pVDZ.

	$\beta_0$	$\mu$	HOMO	LUMO	GAP	BLA
Mol_a	85.07	9.72	−6.25	−1.64	4.53	0.077
Mol_b	229.79	11.31	−6.14	−1.72	4.34	0.086

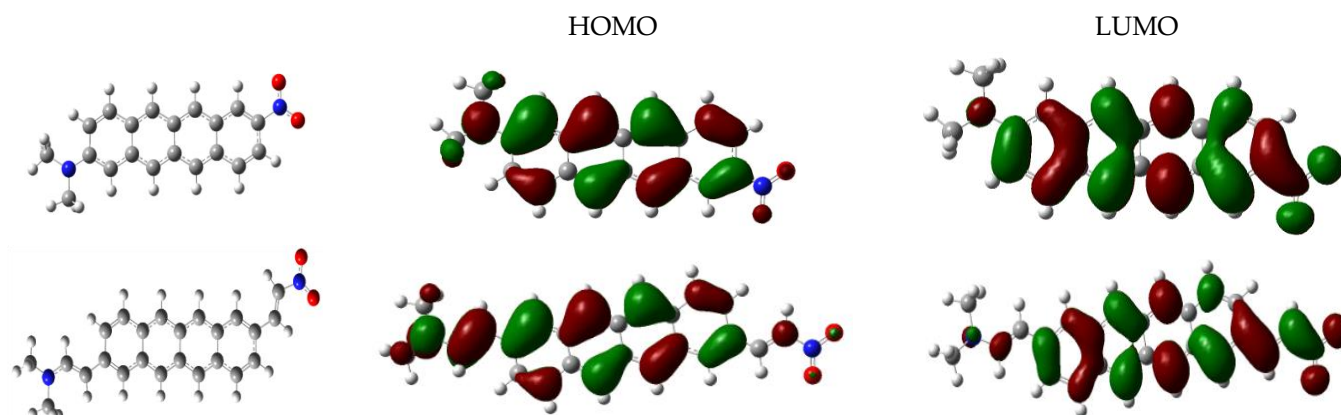
### 3.2. II-Intercalation inside SWCNTs

#### 3.2.1. IIa. Effect of Position

The variation in the different NLO parameters as a function of the position of the paranitroaniline guest molecule inside the carbon nanotube (CNT) is reported in Table 7 and is depicted in Figure 6. According to these results, the charge transfer is largest when the guest molecule is placed in the center of the carbon nanotube, resulting in a maximum value of the static hyperpolarizability. At that position, there is a maximum guest–host interaction. The energy gap hardly varies by varying the position, a finding that is related



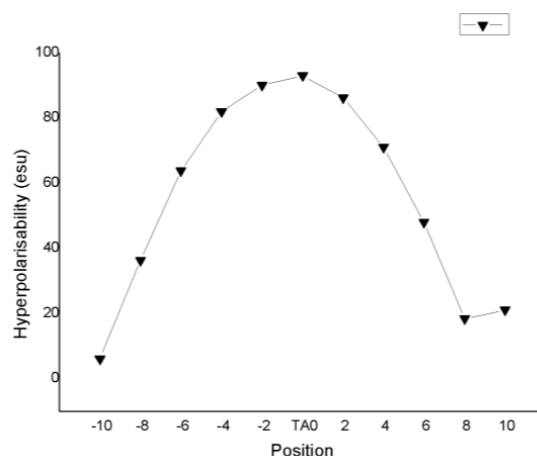
to the fact that the two frontier orbitals HOMO and LUMO are localized mainly on the finite SWCNTs. Finally, the ICT hardly changes with the position of the guest molecule inside the CNT.



**Figure 5.** HOMO and LUMO frontier orbitals of (top) Mol\_a and (bottom) Mol\_b molecules.

**Table 7.** Calculated hyperpolarizability ( $10^{-30}$  esu), dipole moment (Debye), and energy gap (eV) of CNT-PNA, using the CAM-B3LYP/6-31g(d,p) method as a function of the position (in Å) of the PNA inside the SWCNT.

	−10	−8	−6	−4	−2	0	2	4	6	8	10
$\beta_{tot}$	6.14	36.37	63.82	81.97	92.14	93.00	86.28	71.12	47.99	18.46	21.09
$\mu_{tot}$	5.18	3.97	3.09	2.56	2.36	2.33	2.45	2.86	3.72	5.29	7.24
gap	1.85	1.85	1.85	1.85	1.85	1.85	1.85	1.85	1.85	1.85	1.85



**Figure 6.** Evolution of hyperpolarizability ( $10^{-30}$  esu) of CNT-PNA as a function of the PNA position inside the CNT and as obtained with the CAM-B3LYP/6-31g(d,p) method.

### 3.2.2. IIb. Effect of the Nature of the Guest Molecule

In Table 8, we list the values of the first static hyperpolarizability, the dipole moment, and the energy gap for various guests inside the SWCNT. In all cases, the guest is placed at the center of the host. From these results, we observe that the VDA molecule possesses the highest hyperpolarizability despite this not being the largest molecule. The same observation holds true for the dipole moment. These high values are partly due to the longer conjugation because of the vinyl groups on either side of the benzene in the push-pull molecule, as demonstrated in the first part of this study. We also notice that the HOMO–LUMO gap remains constant for the six systems, which, again, can be explained from the localization of those two orbitals to the finite SWCNT.

**Table 8.** Calculated hyperpolarizability ( $10^{-30}$  esu), dipole moment (Debye), and energy gap (eV) of different CNT-guest molecules, using CAM-B3LYP/6-31g(d,p).

	PNA	STIL	TETR	VA	VD	VDA
$\beta_{tot}$	93.00	89.60	68.79	78.82	72.49	114.35
$\mu_{tot}$	2.32	3.46	2.75	2.15	1.97	3.57
gap	1.85	1.85	1.85	1.85	1.85	1.85

### 3.2.3. IIc. Effects Due to Substitution

Finally, we considered the effects of modifying the SWCNT by adding a donor and an acceptor group to its ends (see Figure 3). This was expected to lead to an increase in the charge transfer properties of the whole system. The results are reported in Table 9. Upon substitution,  $\beta_{tot}$  becomes 3 times larger. Similarly, the dipole moment increases significantly, a behavior that is observed in all three spatial directions, x, y, and z. After the substitution, the value of the energy gap decreases only slightly from 1.85 to 1.82 eV.

**Table 9.** Calculated hyperpolarizability ( $10^{-30}$  esu), dipole moment (Debye), and energy gap (eV) of SWCNT-PNA with (denoted TA1) and without (denoted TA0) substitution on the SWCNT, as obtained from the CAM-B3LYP/6-31g(d,p) calculations.

	TA0	TA1
$\beta_x$	−7.92	262.94
$\beta_y$	0.63	−101.40
$\beta_z$	−92.66	−19.93
$\beta_{tot}$	93.00	282.52
$\mu_x$	−2.31	6.78
$\mu_y$	0.09	5.82
$\mu_z$	0.23	0.42
$\mu_{tot}$	2.32	8.94
Gap	1.85	1.82

## 4. Conclusions

The purpose of the present work was to study the effects of functionalization and/or embedding on the NLO properties of some PAHs. Therefore, our focus was not on obtaining very accurate values for specific systems, but on monitoring the changes when modifying the system of interest.

At first, we showed that the functional CAM-B3LYP provided the most accurate description of the properties of interest when using MP2 results as a reference. Furthermore, this was most important for PAHs for which the rings are arranged linearly, as demonstrated in the case of tetracene, a case where long-ranged (exchange) interactions are most pronounced. Moreover, the addition of vinyl groups to the conjugated  $\pi$  bridge led to enhanced NLO responses.

The intercalation of the PAH-derived molecules inside carbon nanotubes also led to increased NLO responses. Finally, the functionalization of the CNT through donor and acceptor groups to the CNT made it possible to increase the intramolecular charge transfer, leading to increased values of the hyperpolarizability and of the dipole moment but, in parallel, an only slightly reduced value of the energy gap.

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